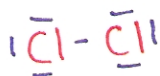
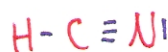
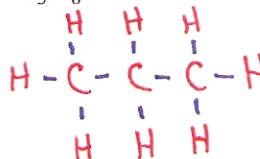


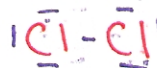
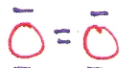
Chemistry Test - Molecular Geometry

2. 1. The polarity of a molecule is a direct result of all of the following EXCEPT
- the presence of polar bonds
 - the shape of the molecule.
 - the types of atoms in the molecule.
 - d. the number of multiple bonds in the molecule.

6. 2. Draw the Lewis structure for the following molecules.

a. Cl_2 b. HCN c. C_3H_8 

6. 3. Which molecule has the greatest bond energy, O_2 , N_2 , Cl_2 ?



4. 4. State the basic assumptions of the VSEPR theory.

Valence Shell Electron Pair Repulsion
electrons repel...

6. 5. Explain how a molecule can have polar bonds but be non-polar.

dipoles are equal and opposite